If there is no stable isotope, the atomic mass of the most stable
isotope is given in parentheses.

Table 4.1 Revised June 2017 by D.E. Groom (LBNL). The atomic number (top left) is the number of protons in the nucleus. The atomic masses (bottom) of stable elements are weighted by isotopic abundances in the Earth’s surface. Atomic masses are relative to the mass of 12C, defined to be exactly 12 u. The exceptions are 1H, 17O, and 18O whose masses are given in natural and commercial samples. Relative isotopic abundances are taken from NIST, https://www.nist.gov/pml/atomic-weights-and-isotopic-compositions-relative-atomic-masses. If there is no stable isotope, the atomic mass of the most stable
isotope is given in parentheses.